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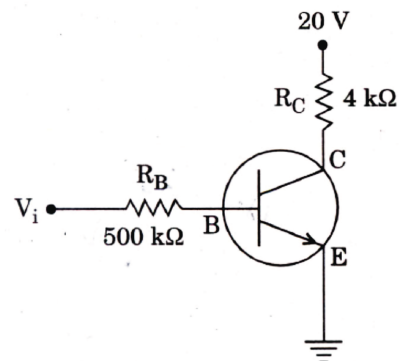
NATIONAL ELIGIBILITY CUM ENTRANCE TEST (UG)

[NEET-2018], **Test held on 06 May 2018**

Subject: PHYSICS

[SET-NN]

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1. The magnetic potential energy stored in a certain inductor 25mJ, when the current in the inductor is 60mA. This inductor is of inductance:

   (A) 13.89 H

   (B) 0.138 H

   (C) 1.389 H

   (D) 138.88 H.

2. An object is placed at a distance of 40 cm from a concave mirror of focal length 15 cm. If the object is displaced through a distance of 20 cm towards the mirror, the displacement of the image will be:

   (A) 36 cm towards the mirror

   (B) 30 cm away from the mirror

   (C) 30 cm towards the mirror

   (D) 36 cm away from the mirror.

3. An em wave is propagating in a medium with a velocity $\vec{V} = V\hat{i}$. The instantaneous oscillating electric field of this em wave is along +y axis. Then the direction of oscillating magnetic field of the em wave will be along:

   (A) $-$ x direction

   (B) $-$ z direction

   (C) $-$ y direction

   (D) + z direction.

4. The refractive index of the material of a prism is $\sqrt{2}$ and the angle of the prism is 30°. One of the two refracting surfaces of the prism is made a mirror inwards, by silver coating. A beam of monochromatic light entering the prism from the other face will retrace its path (after reflection from the silvered surface) if angle of incidence on the prism is :

   (A) zero

   (B) 60°

   (C) 30°

   (D) $45^{o.}$

5. In the circuit shown in the figure, the input voltage V_i is 20 V, $V_{BE} = 0$ and $V_{CE} = 0$. The values of I_B, I_C and β are given by :

   (A) $I_B = 40~\mu A$, $I_C = 5$ mA, $\beta = 125$

   (B) $I_B = 40~\mu A$, $I_C = 10$ mA, $\beta = 250$

   (C) $I_B = 20~\mu A$, $I_C = 5$ mA, $\beta = 250$

   (D) $I_B = 25~\mu A$, $I_C = 5$ mA, $\beta = 200$.

6. In a p-n junction diode, change in temperature due to heating :

   (A) affects the overall V - I characteristics of p-n junction

   (B) affects only reverse resistance

   (C) does not affect resistance of p-n junction

   (D) affects only forward resistance.

7. In the combination of the following gates the output Y can be written in terms of inputs A and B as :
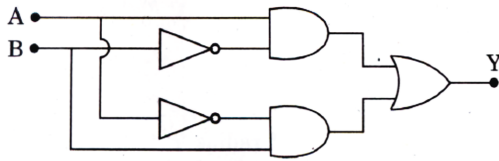
(A) $\overline{A+B}$

(B) $\overline{A \cdot B}$

(C) $\overline{A} \cdot B + A \cdot \overline{B}$

(D) $A \cdot \overline{B} + \overline{A} \cdot B$

8. The power radiated by a black body is P and it radiates maximum energy at wavelength, λ_0. If the temperature of the black body is now changed so that it radiates maximum energy at wavelength $\frac{3}{4}\lambda_0$, the power radiated by it becomes nP. The value of n is :

(A) $\frac{81}{256}$

(B) $\frac{3}{4}$

(C) $\frac{256}{81}$

(D) $\frac{4}{3}$

9. Two wires are made of the same material and have the same volume. The first wire has cross-sectional area A and the second wire has cross-sectional area 3A. If the length of the first wire is increased by $\Delta \ell$ on applying a force F, how much force is needed to stretch the second wire by the same amount?

(A) F

(B) 9F

(C) 4F

(D) 6F.

10. A sample of 0·1 g of water at 100°C and normal pressure (1.013×10^5 Nm^{-2}) requires 54 cal of heat energy to convert to steam at 100°C. If the volume of the steam produced is 167·1 cc, the change in internal energy of the sample, is:

(A) 84.5 J

(B) 104.3 J

(C) 42.2 J

(D) 208.7 J.

11. A small sphere of radius 'r' falls from rest in a viscous liquid. As a result, heat is produced due to viscous force. The rate of production of heat when the sphere attains its terminal velocity, is proportional to :

(A) r^4

(B) r^3

(C) r^5

(D) r^2.

12. When the light of frequency $2\nu_0$ (where ν_0 is threshold frequency), incident on a metal plate, the maximum velocity of electrons emitted is v_1. When the frequency of the incident radiation is increased to $5\nu_0$, the maximum velocity of electrons emitted from the same plate is v_2. The ratio of v_1 to v_2 is :

(A) 2 : 1

(B) 1 : 2

(C) 4 : 1

(D) 1 : 4.

13. For a radioactive material, half-life is 10 minutes. If initially there are 600 number of nuclei, the time taken (in minutes) for the disintegration of 450 nuclei is :

(A) 15

(B) 20

(C) 30

(D) 10.

14. An electron of mass m with an initial velocity $\vec{V} = V_0\hat{i}$ ($V_0 > 0$) enters an electric field $\vec{E} = -E_0\hat{i}$ (E_0 = constant > 0) at t = 0. If λ_0 is its de-Broglie wavelength initially, then its de-Broglie wavelength at time t is:

(A) λ_0

(B) $\dfrac{\lambda_0}{\left(1+\dfrac{eE_0}{mV_0}t\right)}$

(C) $\lambda_0 t$

(D) $\lambda_0\left(1+\dfrac{eE_0}{mV_0}t\right)$.

15. The ratio of kinetic energy to the total energy of an electron in a Bohr orbit of the hydrogen atom, is:

(A) 1 : – 2

(B) 1 : 1

(C) 2 : – 1

(D) 1 : – 1.

16. A tuning fork is used to produce resonance in a glass tube. The length of the air column in this tube can be adjusted by a variable piston. At room temperature of 27°C two successive resonances are produced at 20 cm and 73 cm of column length. If the frequency of the tuning fork is 320 Hz, the velocity of sound in air at 27°C is :

(A) 300 m/s

(B) 330 m/s

(C) 350 m/s

(D) 339 m/s.

17. The electrostatic force between the metal plates of an isolated parallel plate capacitor C having a charge Q and area A, is :

    (A) inversely proportional to the distance between the plates

    (B) independent of the distance between the plates.

    (C) proportional to the square root of the distance between the plates.

    (D) linearly proportional to the distance between the plates.

18. A pendulum is hung from the roof of a sufficiently high building and is moving freely to and fro like a simple harmonic oscillator. The acceleration of the bob of the pendulum is 20 m/s^2 at a distance of 5 m from the mean position. The time period of oscillation is :

    (A) 1 s   (B) 2π s

    (C) 2 s   (D) π s.

19. An electron falls from rest through a vertical distance h in a uniform and vertically upward directed electric field E. The direction of electric field is now reversed, keeping its magnitude the same. A proton is allowed to fall from rest in it through the same vertical distance h. The time of fall of the electron, in comparison to the time of fall of the proton is :

    (A) equal

    (B) smaller

    (C) 10 times greater

    (D) 5 times greater.

20. The kinetic energies of a planet in an elliptical orbit about the Sun, at positions A, B and Care K_A, K_B and K_C, respectively. AC is the major axis and SB is perpendicular to AC at the position of the Sun S as shown in the figure. Then :
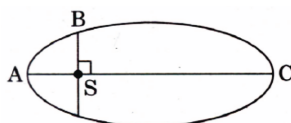
    (A) $K_B > K_A > K_C$ (B) $K_A < K_B < K_C$

    (C) $K_B < K_A < K_C$ (D) $K_A > K_B > K_C$.

21. A solid sphere is in rolling motion. In rolling motion a body possesses translational kinetic energy (K_t) as well as rotational kinetic energy (K_r) simultaneously. The ratio $K_t : (K_t + K_r)$ for the sphere is :

    (A) 2 : 5   (B) 7 : 10

    (C) 10 : 7   (D) 5 : 7.

22. If the mass of the Sun were ten times smaller and the universal gravitational constant were ten times larger in magnitude, which of the following is not correct?

    (A) 'g' on the Earth will not change.

    (B) Raindrops will fall faster.

    (C) Time period of a simple pendulum on the Earth would decrease

    (D) Walking on the ground would become more difficult.

23. A solid sphere is rotating freely about its symmetry axis in free space. The radius of the sphere is increased keeping its mass same. Which of the following physical quantities would remain constant for the sphere?

    (A) Angular momentum

    (B) Angular velocity

    (C) Rotational kinetic energy

    (D) Moment of inertia.

24. A metallic rod of mass per unit length 0·5 kg m^{-1} is lying horizontally on a smooth inclined plane which makes an angle of 30° with the horizontal. The rod is not allowed to slide down by flowing a current through it when a magnetic field of induction 0·25 T is acting on it in the vertical direction. The current flowing in the rod to keep it stationary is:

    (A) 11.32 A   (B) 7.14 A

    (C) 14.76 A   (D) 5.98 A.

25. An inductor 20 mH, a capacitor 100 μF and a resistor 50 Ω are connected in series across a source of emf, V = 10 sin 314 t. The power loss in the circuit is :

   (A) 1.13 W   (B) 0.79 W

   (C) 2.74 W   (D) 0.43 W.

26. A thin diamagnetic rod is placed vertically between the poles of an electromagnet. When the current in the electromagnet is switched on, then the diamagnetic rod is pushed up, out of the horizontal magnetic field. Hence the rod gains gravitational potential energy. The work required to do this comes from:

   (A) the induced electric field due to the changing magnetic field

   (B) the current source

   (C) the lattice structure of the material of the rod

   (D) the magnetic field.

27. Current sensitivity of a moving coil galvanometer is 5 div/mA and its voltage sensitivity (angular deflection per unit voltage applied) is 20 div/V. The resistance of the galvanometer is:

   (A) 500 Ω (B) 40 Ω

   (C) 250 Ω (D) 25 Ω.

28. The volume (V) of a monatomic gas varies with its temperature (T), as shown in the graph. The ratio of work done by the gas, to the heat absorbed by it, when it undergoes a change from state A to state B, is:
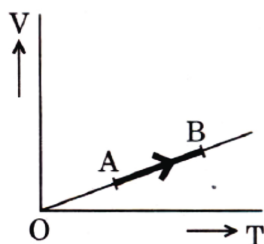
   (A) $\dfrac{2}{7}$ (B) $\dfrac{2}{5}$

   (C) $\dfrac{1}{3}$ (D) $\dfrac{2}{3}$.

29. The fundamental frequency in an open organ pipe is equal to the third harmonic of a closed organ pipe. If the length of the closed organ pipe is 20 cm, the length of the open organ pipe is:

   (A) 16 cm   (B) 13.2 cm

   (C) 12.5 cm   (D) 8 cm.

30. The efficiency of an ideal heat engine working between the freezing point and boiling point of water, is:

   (A) 12.5%   (B) 26.8%

   (C) 6.25%   (D) 20%.

31. At what temperature will the rms speed of oxygen molecules become just sufficient for escaping from the Earth's atmosphere?

   (Given: Mass of oxygen molecule (m) = 2·76 × 10^{-26} kg Boltzmann's constant k_B = 1·38 × 10^{-23} J K^{-1})

   (A) 1.254×10^4 K

   (B) 2.508×10^4 K

   (C) 5.016×10^4 K

   (D) 8.360×10^4 K.

32. Unpolarised light is incident from air on a plane surface of a material of refractive index 'μ'. At a particular angle of incidence 'i', it is found that the reflected and refracted rays are perpendicular to each other. Which of the following options is correct for this situation?

   (A) $i = \tan^{-1}\left(\dfrac{1}{\mu}\right)$

   (B) Reflected light is polarised with its electric vector parallel to the plane of incidence

   (C) $i = \sin^{-1}\left(\dfrac{1}{\mu}\right)$

   (D) Reflected light is polarised with its electric vector perpendicular to the plane of incidence.

33. In Young's double slit experiment the separation d between the slits is 2 mm, the wavelength λ of the light used is 5896 Å and distance D between the screen and slits is 100 cm. It is found that the angular width of the fringes is 0.20°. To increase the fringe angular width to 0.21° (with same λ and D) the separation between the slits needs to be changed to :

    (A) 1.7 mm   (B) 1.8 mm

    (C) 2.1 mm   (D) 1.9 mm.

34. An astronomical refracting telescope will have large angular magnification and high angular resolution, when it has an objective lens of :

    (A) small focal length and small diameter

    (B) small focal length and large diameter

    (C) large focal length and large diameter

    (D) large focal length and small diameter.

35. A body initially at rest and sliding along a frictionless track from a height h (as shown in the figure) just completes a vertical circle of diameter AB = D. The height h is equal to :
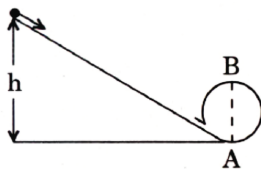
    (A) $\frac{5}{4}D$ (B) $\frac{3}{2}D$

    (C) $\frac{7}{5}D$ (D) D.

36. Three objects, A : (a solid sphere), B : (a thin circular disk) and C : (a circular ring), each have the same mass M and radius R. They all spin with the same angular speed m about their own symmetry axes. The amounts of work (W) required to bring them to rest, would satisfy the relation:

    (A) $W_A > W_C > W_B$

    (B) $W_C > W_B > W_A$

    (C) $W_B > W_A > W_C$

    (D) $W_A > W_B > W_C$.

37. Which one of the following statements is incorrect?

    (A) Coefficient of sliding friction has dimensions of length.

    (B) Rolling friction is smaller than sliding friction.

    (C) Frictional force opposes the relative motion

    (D) Limiting value of static friction is directly proportional to normal reaction.

38. A moving block having mass m, collides with another stationary block having mass 4m. The lighter block comes to rest after collision. When the initial velocity of the lighter block is v, then the value of coefficient of restitution (e) will be :

    (A) 0.4   (B) 0.5

    (C) 0.8   (D) 0.25.

39. A carbon resistor of (47 ± 4.7) kΩ is to be marked with rings of different colours for its identification. The colour code sequence will be :

    (A) Green – Orange – Violet – Gold

    (B) Violet – Yellow – Orange – Silver

    (C) Yellow – Green – Violet – Gold

    (D) Yellow – Violet – Orange – Silver.

40. A set of 'n' equal resistors, of value 'R' each, are connected in series to a battery of emf 'E' and internal resistance 'R'. The current drawn is I. Now, the 'n' resistors are connected in parallel to the same battery. Then the current drawn from battery becomes 10 I. The value of 'n' is:

    (A) 9

    (B) 10

    (C) 20

    (D) 11.

41. A battery consists of a variable number 'n' of identical cells (having internal resistance 'r' each) which are connected in series. The terminals of the battery are short-circuited and the current I is measured. Which of the graphs shows the correct relationship between I and n?

(A)

(B)

(C)

(D)
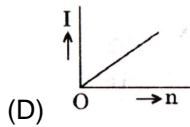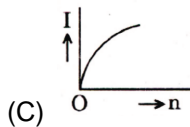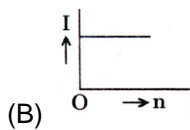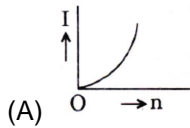
42. A toy car with charge q moves on a-frictionless horizontal plane surface under the influence of a uniform electric field $\vec{E}$. Due to the force $q\vec{E}$, its velocity increases from 0 to 6 m/s in one second duration. At that instant the direction of the field is reversed. The car continues to move for two more seconds under the influence of this field. The average velocity and the average speed of the toy car between 0 to 3 seconds are respectively:

(A) 1.5 m/s, 3m/s

(B) 2 m/s, 4 m/s

(C) 1 m/s, 3.5 m/s

(D) 1 m/s, 3 m/s.

43. A block of mass m is placed on a smooth inclined wedge ABC of inclination θ as shown in the figure. The wedge is given an acceleration 'a' towards the right. The relation between a and θ for the block to remain stationary on the wedge is:

(A) a = g tan θ

(B) $a = \dfrac{g}{\cosec \theta}$

(C) a = g cos θ
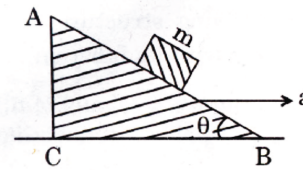
(D) $a = \dfrac{g}{\sin\theta}$.

44. A student measured the diameter of a small steel ball using a screw gauge of least count 0·001 cm. The main scale reading is 5 mm and zero of circular scale division coincides with 25 divisions above the reference level. If screw gauge has a zero error of –0·004 cm, the correct diameter of the ball is

(A) 0.529 cm

(B) 0.521 cm

(C) 0.053 cm

(D) 0.525 cm.

45. The moment of the force, $\vec{F} = 4\hat{i} + 5\hat{j} - 6\hat{k}$ at (2, 0, –3), about the point (2, –2, –2), is given by:

(A) $-7\hat{i} - 4\hat{j} - 8\hat{k}$

(B) $-8\hat{i} - 4\hat{j} - 7\hat{k}$

(C) $-7\hat{i} - 8\hat{j} - 4\hat{k}$

(D) $-4\hat{i} - \hat{j} - 8\hat{k}$.

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NATIONAL ELIGIBILITY CUM ENTRANCE TEST (UG)
[NEET-2018], Test held on 06 May 2018
Subject :BIOLOGY
[SET-NN]

46. Which of the following hormones can play a significant role in osteoporosis:

    (A) Parathyroid hormone and Prolactin   (B) Aldosterone and Prolactin

    (C) Estrogen and Parathyroid hormone   (D) Progesterone and Aldosterone

47. Which of the following is an amino acid derived hormone :

    (A) Estriol   (B) Epinephrine

    (C) Estradiol   (D) Ecdysone

48. Which of the following structures of regions is incorrectly paired with its function :

    (A) Corpus callosum   :   band of fibers connecting left and right cerebral hemisphere

    (B) Medulla oblongata   :   controls respiration and cardiovascular

    (C) Hypothalamus   :   Production of releasing hormones and regulation of temperature hunger and thirst

    (D) Limbic system   :   consists of fibre tracts that interconnect different regions of brain; controls movement

49. The transparent lens in the human eye is held in its place by :

    (A) smooth muscles attached to the ciliary body   (B) ligaments attached to the ciliary body

    (C) smooth muscles attached to the iris   (D) ligaments attached to the iris

50. The amnion of mammalian embryo is derived from

    (A) ectoderm and endoderm

    (B) ectoderm and mesoderm

    (C) mesoderm and trophoblast

    (D) endoderm and mesoderm

51. Hormones secreted by the placenta to maintain pregnaryare :

    (A) hCG progestogens, estrogens glucocorticoids

    (B) hCG, hPL, progestogens, prolactin

    (C) hCG, hpt, progestogens, estrogens

    (D) hCG ,hPL, estrogens, relaxin, oxytocin

52. The contraceptive 'SAHELI' :

    (A) is a post – coital contraceptive

    (B) blocks estrogen receptors in the uterus, preventing eggs from getting implanted

    (C) is an IUD

    (D) increase the concentration of estrogen and prevents ovulation in female

53. The difference between spermiogenesis and spermiationis :

    (A)In spermiogenesis spermatozoa are formed, while in spermiation spermatozoa are released from sertoli cells into the cavity of seminiferous tubules.

    (B) In spermiogenesis spermatids are formed, while in spermiation spermatozoa are formed.

    (C) In spermiogenesis spermatozoa from sertoli cells are released into the cavity of seminiferous tubules, while in spermiationspermatozoa are formed.

    (D) In spermiogenesis spermatozoa are formed, while Inspermiation spermatids are formed.

54. A woman has an X – linked condition on one of her chromosomes. This chromosome can be inherited by

    (A) Both sons and daughters   (B) Only daughters

    (C) Only grandchildren   (D) Only sons

55. According to Hugo de Vries, the mechanism of evolution is :

    (A) Minor mutations   (B) Multiples step mutations

    (C) Phenotypic variations   (D) Saltation

56. All of the following are part of an operon except :

    (A) a promoter   (B) an operator

    (C) an enhancer   (D) structural genes

57. AGGTATCGCAT is a sequence from the coding strand of a gene. What will be the corresponding sequence of the transcribed mRNA :

    (A) UCCAUAGCGUA   (B) AGGUAUCGCAU

    (C) ACCUAUGCGAU   (D) UGGTUTCGCAT

58. Match the items given in Column I with those in Column II and select the correct option given below :

Column –I		Column –II	
a.	Proliferative Phase	i.	Breakdown of endometrial lining
b.	Secretory Phase	ii.	Follicular Phase
c.	Menstruation	iii.	Luteal Phase

    (A) a – iii, b – i , c - ii   (B) a – iii, b – ii , c - i

    (C) a – ii, b – iii , c –I   (D) a – i, b – iii , c - ii

59. In which disease does mosquito transmitted pathogen cause chronic inflammation of lymphatic vessels:

    (A) Amoebiasis   (B) Elephantiasis

    (C) Ringworm disease   (D) Ascariasis

60. Among the following sets of examples for divergent evolution, select the incorrect option :

    (A) Eye of octopus bat and cheetah
    (B) Forelimbs of man, bat and cheetah

    (C) Brain of bat, man and cheetah
    (D) Heart of bat, man and cheetah

61. The similarity of bone structure in the forelimbs of many vertebrates is an example of :

    (A) Adaptive radiation
    (B) Homology

    (C) Convergent evolution
    (D) Analogy

62. Which of the following is not an autoimmune disease :

    (A) Vitiligo
    (B) Psoriasis

    (C) Alzheimer's disease
    (D) Rheumatoid arthritis

63. Which of the following characteristics represent 'Inheritance of blood groups' in humans :

    a. Dominance
    b. Co- dominance

    b. Multiple allele
    c. Incompletedominace

    e. Polygenic inheritance

    (A) a, c and e
    (B) b, c and e

    (C) b, d and e
    (D) a, b and c

64. Conversion of milk to curd improves its nutritional value by increasing the amount of :

    (A) Vitamin E
    (B) Vitamin D

    (C) Vitamin B_{12}
    (D) Vitamin A

65. Match the items given in Column I with those in Column II and select the correct option given below :

	Column – I		Column – II
a.	Eutrophication	i.	UV – B radiation
b.	Sanitary landfill	ii.	Deforestation
c.	Snow blindness	iii.	Nutrient enrichment
d.	Jhum cultivation	iv.	Waste disposal

    (A) a – i , b –ii, c- iv, d – iii
    (B) a – ii , b – i, c- iii, d - iv

    (C) a – iii, b – iv, c- i, d – ii
    (D) a – i, b – iii, c- iv, d – ii

66. Which one of the following population interactions is widely used in medical science for the production of antibiotics :

    (A) Amensalism
    (B) Commensalism

    (C) Parasitism
    (D) Mutualism

67. All of the following are included in 'Ex- situ conservation' except :

    (A) Seed banks

    (B) Wildlife safari parks

    (C) Botanical gardens

(D) Sacred groves

68. In a growing population of a country :

(A) pre-reproductive individuals are less than the reproductive individuals

(B) pre- reproductive individuals are more than the reproductive individuals

(C) reproductive and pre – reproductive individuals are equal in number

(D) reproductive individuals are less than the post – reproductive individuals

69. Which part of poppy plant is used to obtain the drug "Smack" :

(A) Leaves   (B) Flowers

(C) Roots   (D) Latex

70. Which of the following gastric cells indirectly help in erythropoiesis :

(A) Parietal cells   (B) Chief cells

(C) Goblet cells   (D) Mucous cells

71. Match the items given in Column I with those in Column –II and select the correct option given below :

Column –I   Column –II

a. Fibrinogen   i.   Osmotic balance

b. globulin   ii.   Blood clotting

c. Albumin   iii. Defence mechanism

(A) a – ii, b – iii , c – I   (B) a – iii, b – ii , c - i

(C) a – i, b – iii, c – ii   (D) a – i, b – ii , c – iii

72. Which of the following is an occupational respiratory disorder :

(A) Emphysema   (B) Anthracis

(C) Botulism   (D) Silicosis

73. Calcium is important in skeletal muscle contraction because it :

(A) prevents the formation of bonds between the myosin cross bridges and the actin filament

(B) binds to troponin to remove the masking of active sites on actin for myosin

(C) detaches the myosin head from the actin filament

(D) activates the myosin ATPase by binding to it

74. Nissl bodies are mainly composed of :

(A) Free ribosomes and RER

(B) Proteins and lipids

(C) Nucleic acids and SER

(D) DNA and RNA

75. Which of these statements is incorrect :

(A) Oxidative phosphorylation takes place in outer mitochondrial membrane

(B) Enzymes of TCA cycle are present in mitochondrial matrix

(C) Glycolysis operates as long as it is supplied with NAD that can pick up hydrogen atoms

(D) Glycolysis occurs in cytosol

76. Select the incorrect match :

    (A) Polytene chromosomes   -   Oocytes of amphibian

    (B) Lampbrush chromosomes   -   Diplotene bivalents

    (C) Submetacentric chromosomes   -   L-shaped chromosomes

    (D) Allosomes   -   Sex chromosomes

77. Which of the following terms describe human dentition :

    (A) Pleurodont, Diphyodont, Heterodont

    (B) Thecodont, Diphyodont , Homodont

    (C) Pleurodont, Monophyodont, Homodont

    (D) Thecodont, Diphyodont, Heterodont

78. Which of the following events does not occur in rough endoplasmic reticulum :

    (A) Phospholipid synthesis   (B) Protein folding

    (C) Cleavage of signal peptide   (D) Protein glycosylation

79. Many ribosomes may associate with a single mRNa to form form multiple copies of a polypeptide simultaneously. Such strings of ribosomes are termed as :

    (A) Nucleosome   (B) Polysome

    (C) Plastidome   (D) Polyhedral bodies

80. Ciliates differ from all other protozoans in :

    (A) having two types of nuclei   (B) using flagella for locomotion

    (C) using pseudopodia for capturing prey   (D) having a contractile vacuole for removing

81. Identify the vertebrate group of animals characterized by crop and gizzard in its digestive system :

    (A) Osteichthyes   (B) Amphibia

    (C) Aves   (D) Reptilia

82. Which one of these animals is not a homeotherm :

    (A) Psittacula   (B) Macropus

    (C) Camelus   (D) Chelone

83. Which of the following feature is used to identify a male cockroach from a female cockroach :

    (A) Presence of anal cerci

    (B) Presence of a boat shaped sternum on the 9th abdominal segment

    (C) Forewings with darker tegmina

    (D) Presence of caudal styles

84. Which of the following animals does not undergo metamorphosis :

    (A) Starfish   (B) Earthworm

    (C) Moth   (D) Tunicate

85. Which of the following organisms are known as chief producers in the oceans :

    (A) Euglenoids   (B) Dinoflagellates

    (C) Cyanobacteria   (D) Diatoms

86. Which of the following options correctly represents the lung conditions in asthma and emphysema, respectively :

    (A) Decreased respiratory surface; Inflammation of bronchioles

    (B) Inflammation of bronchioles Decreased respiratory surface

    (C) Increased respiratory surface; Inflammation of bronchioles

    (D) Increased number of bronchioles, Increased respiratory surface

87. Match the items given in Column – I with those in Column I I and select the correct option given below :

Column - I		Column – II	
a.	Tricuspid valve	i.	Between left atrium and left ventricle
b.	Bicuspid valve	ii.	Between right ventricle and pulmonary artery
c.	Semilunar valve	iii.	Between right atrium and right ventricle

    (A) a – ii , b – i, c – iii   (B) a – iii , b – i, c - ii

    (C) a – i , b – ii, c – iii   (D) a – i , b – iii, c – ii

88. Match the items given in Column I with those in Column II and select the correct option given below :

Column –I		Column –II	
a.	Tidal volume	i.	2500 – 3000 mL
b.	Inspiratory Reserve volume	ii.	1100 - 1200 mL
c.	Expiratory Reserve volume	iii.	500 – 550 mL
d.	Residual volumne	iv.	1000 – 1100 mL

    (A) a – iv, b – iii, c – ii, d –I   (B) a – iii, b – ii, c – i, d - iv

    (C) a – i, b – iv, c – ii, d – iii   (D) a – iii, b – i, c – iv, d – ii

89. Match the items given in Column I with those in Column II and select the correct option given below :

Column – I		Column - II	
a.	Glycosuria	i.	Accumulation of uric acid in joints
b.	Gout	ii.	Mass of crystallised salts within the kidney
c.	Renal calculi	iii	Inflammation in glomeruli
d.	Glomerular nephritis	iv.	Presence of glucose in urine

    (A) a – iv, b – i, c – ii, d – iii

    (B) a – iii, b – ii, c – iv, d - i

    (C) a – ii, b – iii, c – i, d – iv

    (D) a – i, b – ii, c – iii, d – iv

90. Match the items given in Column I with those in Column II and select the correct option given below :

| Column – I | Column –II |
| (Function) | (Part of Excretory system) |

a.   Ultrafiltration   i.   Henle's loop

b.   Concentration of urine   ii.   Ureter

c.   Transport of urine   iii.   Urinary bladder

d.   Storage of urine   iv.   Malpighian corpuscle

                                                             v.   Proximal convoluted tubule

(A) a – v, b – iv, c – i, d – iii   (B) a – iv, b – v, c – ii, d - iii

(C) a – v, b – iv, c – i, d – ii   (D) a – iv, b – i, c – ii, d – iii

91. Secondary xylem and phloem in dicot stem areproduced by

(A) Axillary meristems   (B) Apical meristems

(C) Phellogen   (D) Vascular cambium

92. Pneumatophores occur in

(A) Submerged hydrophytes   (B) Halophytes

(C) Carnivorous plants   (D) Free-floating hydrophytes

93. Plants having little or no secondary growth are

(A) Cycads   (B) Grasses

(C) Conifers   (D) Deciduous angiosperms

94. Select the wrong statement

(A) Mitochondria are the powerhouse of the cell in all kingdoms except Monera

(B) Cell wall is present in members of Fungi and Plantae

(C) Pseudopodia are locomotory and feeding structures in Sporozoans

(D) Mushrooms belong to Basidiomycetes

95. Casparian strips occur in

(A) Endodermis   (B) Epidermis

(C) Cortex   (D) Pericycle

96. Sweet potato is a modified

(A) Rhizome   (B) Stem

(C) Tap root   (D) Adventitious root

97. Which of the following statements is correct ?

(A) Stems are usually unbranched in both Cycas and Cedrus

(B) Ovules are not enclosed by ovary wall in gymnosperms

(C) Horsetails are gymnosperms

(D) Selaginella is heterosporous, while Salvinia is homosporous

98. What type of ecological pyramid would be obtained with the following data ?

Secondary consumer : 120 g

Primary consumer : 60 g

Primary producer : 10 g

(A) Upright pyramid of biomass   (B) Inverted pyramid of biomass

(C) Upright pyramid of numbers   (D) Pyramid of energy

99. World Ozone Day is celebrated on

(A) 22nd April   (B) 5th June

(C) 16th September   (D) 21st April

100. In stratosphere, which of the following elements acts as a catalyst in degradation of ozone and release of molecular oxygen ?

(A) Oxygen   (B) Carbon

(C) Fe   (D) Cl

101. Natality refers to

(A) Number of individuals entering a habitat   (B) Death rate

(C) Number of individuals leaving the habitat   (D) Birth rate

102. Niche is

(A) The functional role played by the organism where it lives

(B) All the biological factors in the organism's environment

(C) The range of temperature that the organism needs to live

(D) The physical space where an organism lives

103. Which of the following is a secondary pollutant ?

(A) O_3 (B) CO

(C) SO_2 (D) CO_2

104. Winged pollen grains are present in

(A) Pinus   (B) Mustard

(C) Mango   (D) Cycas

105. Afterkaryogamy followed by meiosis, spores are produced exogenously in

(A) Saccharomyces   (B) Neurospora

(C) Agaricus   (D) Alternaria

106. Which one is wrongly matched ?

(A) Unicellular organism – Chlorella

(B) Uniflagellategamets – Polysiphonia

(C) Gemma cups – Marchantia

(D) Biflagellate zoospores – Brown algae

107. Match the items given in Column I with those in Column II and select the correct option given below :

	Column – I		Column – II
(a)	Herbarium	(i)	It is a place having a collection of preserved plants and animal
(b)	Key	(ii)	A list that enumerates methodically all the species found in an area With brief description aiding identification.
(c)	Museum	(iii)	Is a place where dried and pressed plant specimens mounted on Sheets are kept.
(d)	Catalogue	(iv)	A booklet containing a list of characters and their alternates which Are helpful in identification of various taxa.

(A) a – iii, b – iv, c – i, d – ii   (B) a – i, b – iv, c – iii, d - ii

(C) a – ii, b – iv, c – iii, d – i   (D) a – iii, b – ii, c – i, d – iv

108. Which one of the following plants shows a very close relationship with a species of moth, where none of the two can complete its life cycle without the other?

(A) Viola   (B) Hydrilla

(C) Banana   (D) Yucca

109. Pollen grains can be stored for several years in liquid nitrogen having a temperature of

(A) -160^0C   (B) -120^0C

(C) -196^0C   (D) -80^0C

110. In which of the following forms is iron absorbed by plants ?

(A) Both ferric and ferrous   (B) Ferric

(C) Free element   (D) Ferrous

111. Which of the following elements is responsible for maintaining turgor in cells ?

(A) Calcium   (B) Magnesium

(C) Potassium   (D) Sodium

112. Double fertilization is

(A) Syngamy and triple fusion

(B) Fusion of two male gametes of a pollen tube with two different eggs

(C) Fusion of two male gametes with one egg

(D) Fusion of one male gamete with two polar nuclei

113. What is the role of NAD^+ in cellular respiration ?

(A) It is the final electron acceptor for anaerobic respiration

(B) It functions as an enzyme

(C) It is a nucleotide source for ATP synthesis

(D) It functions as an electron carrier

114. Oxygen is not produced during photosynthesis by

    (A) Chara
    (B) Green sulphur bacteria
    (C) Cycas
    (D) Nostoc

115. The Golgi complex participates in

    (A) Activation of amino acid
    (B) Fatty acid breakdown
    (C) Respiration in bacteria
    (D) Formation of secretory vesicles

116. Stomatal movement is not affected by

    (A) CO_2 concentration
    (B) Temperature
    (C) O_2 concentration
    (D) Light

117. Stomata in grass leaf are

    (A) Barrel shaped
    (B) Dumb-bell shaped
    (C) Rectangular
    (D) Kidney shaped

118. Which of the following is true for nucleolus ?

    (A) It is a site for active ribosomal RNA synthesis

    (B) Larger nucleoli are present in dividing cells

    (C) It takes part in spindle formation

    (D) It is a membrane – bound structure

119. Which of the following is not a product of light reaction of photosynthesis ?

    (A) Oxygen
    (B) ATP
    (C) NADPH
    (D) NADH

120. The stage during which separation of the paired homologous chromosomes begins is

    (A) Zygotene
    (B) Pachytene
    (C) Diakinesis
    (D) Diplotene

121. The two functional groups characteristic of sugars are

    (A) Carbonyl and hydroxyl
    (B) Hydroxyl and methyl
    (C) Carbonyl and phosphate
    (D) Carbonyl and methyl

122. Which among the following is not a prokaryote ?

    (A) Oscillatoria
    (B) Saccharomyces
    (C) Nostoc
    (D) Mycobacterium

123. Offsets are produced by

    (A) Parthenogenesis

    (B) Meiotic divisions

    (C) Parthenocarpy

    (D) Mitotic divisions

124. Select the correct statement :

(A) Transduction was discovered by S. Altman

(B) Franklin Stahl coined the term "linkage".

(C) Spliceosomes take part in translation

(D) Punnett square was developed by a British scientist

125. Which of the following has proved helpful in preserving pollen as fossils ?

(A) Sporopollenin   (B) Pollenkitt

(C) Oil content   (D) Cellulosic intine

126. Which of the following pairs is wrongly matched ?

(A) T.H. Morgan   :   Linkage

(B) Starch synthesis in pea   :   Multiple alleles

(C) XO type sex determination :   Grasshopper

(D) ABO blood grouping   :   Co-dominance

127. Select the correct match :

(A) Francois Jacob and Jacques Monod   –   Lac operon

(B) Alec Jeffreys   –   Streptococcus pneumoniae

(C) Matthew Meselson and F. Stahl   –   Pisumsativum

(D) Alfred Hershey and Martha Chase   –   TMV

128. Which of the following flowers only once in its life-time ?

(A) Papaya   (B) Bamboo species

(C) Mango   (D) Jackfruit

129. The experimental proof for semiconservative replication of DNA was first shown in a

(A) Virus   (B) Fungus

(C) Plant   (D) Bacterium

130. Which of the following is commonly used as a vector for introducing a DNA fragment in human lymphocytes ?

(A) pBR 322   (B) Retrovirus

(C) λ phase   (D) Ti plasmid

131. Select the correct match :

(A) G. Mendel   –   Transformation

(B) Ribozyme   –   Nucleic acid

(C) T.H. Morgan   –   Transduction

(D) F_2 × Recessive parent   –   Dihybrid cross

132. Use of bioresources by multinational companies and organizations without authorization from the concerned country and its people is called

    (A) Bioexploitation
    (B) Bio-infringement
    (C) Biodegradation
    (D) Biopiracy

133. A 'new' variety of rice was patented by a foreign company, though such varieties have been present in India for a long time. This is related to

    (A) Basmati
    (B) Co-667
    (C) Lerma Rojo
    (D) Sharbati Sonora

134. In India, the organization responsible for assessing the safety of introducing genetically modified organisms for public use is

    (A) Genetic Engineering Appraisal Committee (GEAC)

    (B) Indian Council of Medical Research (ICMR)

    (C) Research Committee on Genetic Manipulation (RCGM)

    (D) Council for Scientific and Industrial Research (CSIR)

135. The correct order of steps in Polymerase Chain Reaction (PCR) is

    (A) Denaturation, Annealing, Extension
    (B) Extension, Denaturation, Annealing
    (C) Denaturation, Extension, Annealing
    (D) Annealing, Extension, Denaturation

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[SET - NN]

136. Identify the major products P, Q and R in the following sequence of reactions :
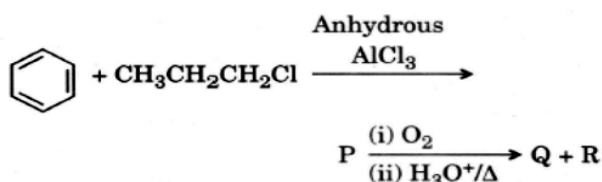
137. Which of the following compounds can form a zwitterion :

    (A) Glycine

    (B) Aniline

    (C) Benzoicv acid

    (D) Acetanilide

138. The geometry and magnetic behaviour of the complex $[Ni(CO)_4]$ are

   (A)   tetrahedral geometry and paramagnetic

   (B)   square planar geometry and diamagnetic

   (C)   square planar geometry and paramagnetic

   (D)   square planar geometry and paramagnetic

139. Iron carbonyl, $Fe(CO)_5$ is

   (A)   dinuclear   (B)   tetranuclear

   (C)   trinuclear   (D)   mononuclear

140. Match the metal ions given in Column I with the spin magnetic moments of the ions given in Column II and assign the correct code :

   Column I   Column II

   a. Co^{3+} i. $\sqrt{8}$ B. M.

   b. Cr^{3+} ii. $\sqrt{35}$ B. M.

   c. Fe^{3+} iii. $\sqrt{3}$ B. M.

   d. Ni^{2+} iv. $\sqrt{24}$ B. M.

                                       iv. $\sqrt{15}$ B. M.

   (A)   a-iii, b-v, c-i. d-ii   (B)   a-iv, b-v, c-ii, d-i

   (C)   a-iv, b-i, c-ii, d-iii   (D)   a-i, b-ii, c-iii, d-iv

141. The type of isomerism shown by the complex $[CoCl_2(en)_2]$ is

   (A)   Linkage isomerism   (B)   Geometrical isomerism

   (C)   Ionization isomerism (D)   Coordination isomerism

142. Which one of the following ions exhibits d-d transition and paramagnetism as well :

   (A) MnO_4^{2-} (B) CrO_4^{2-}

   (C) MnO_4^{-} (D) $Cr2O_7^{2-}$

143. Which of the following molecules represents the order of hybridisation sp^2, sp^2, sp, sp from left to right atoms :

   (A) $CH_3—CH=CH—CH_3$ (B) $HC \equiv C - C \equiv CH_2$

   (C) $CH_2=CH—CH=CH_2$ (D) $CH_2 = CH - C \equiv CH$

144. Which of the following carbocations is expected to be most stable

(A)

(B)

(C)

(D)
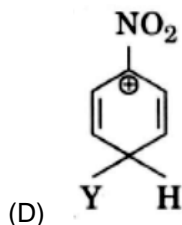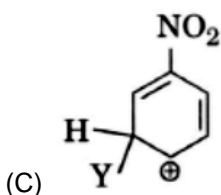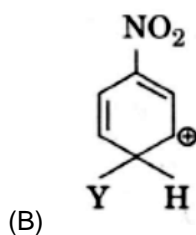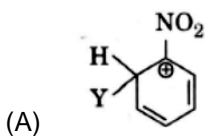
145. Which of the following is correct with respect to --I effect of the substituents (R = alkyl) :

   (A) $-NR_2 > -OR > -F$ (B) $-NH_2 < -OR < -F$

   (C) $-NH_2 > -OR > -F$ (D) $-NR_2 < -OR < -F$

146. The solubility of $BaSO_4$ in water is 2.42×10^{-3} g L^{-1} at 298 K. The value of its solubility product (K_{sp}) will be (Given molar mass of $BaSO_4 = 233$ g mol^{-1})

   (A) 1.08×10^{-8} mol^2 L^{-2} (B) 1.08×10^{-10} mol^2 L^{-2}

   (C) 1.08×10^{-14} mol^2 L^{-2} (D) 1.08×10^{-12} mol^2 L^{-2}

147. Given van der Waals constant for NH_3, H_2, O_2 and CO_2 are respectively 4.17, 0.244, 1.36 and 3.59, which one of the following gases is most easily liquefied :

   (A) CO_2 (B) NH_3

   (C) O_2 (D) H_2

148. Following solutions were prepared by mixing different volumes of NaOH and HCl of different concentrations:

   a. 60 mL $\dfrac{M}{10}$ HCl + 40 mL $\dfrac{M}{10}$ NaOH   b. 55 mL $\dfrac{M}{10}$ HCl + 45 mL $\dfrac{M}{10}$ NaOH

   c. 75 mL $\dfrac{M}{5}$ HCl + 25 mL $\dfrac{M}{5}$ NaOH   d. 100 mL $\dfrac{M}{10}$ HCl + 100 mL $\dfrac{M}{10}$ NaOH

   pH of which one of them will be equal to 1 :

   (A) c   (B) b

   (C) d   (D) a

149. On which of the following properties does the coagulating power of an ion depend :

   (A) The sign of charge on the ion alone

   (B) The magnitude of the charge on the ion alone

   (C) Both magnitude and sign of the charge on the ion

   (D) Size of the ion alone

150. Which of the following statements is not true for halogens :

   (A) Chlorine has the highest electron-gain enthalpy.

   (B) All form monobasic oxyacids.

   (C) All but fluorine show positive oxidation states.

   (D) All are oxidizing agents.

151. Considering Ellingham diagram, which of the following metals can be used to reduce alumina :

   (A) Cu   (B) Fe

   (C) Mg   (D) Zn

152. The correct order of atomic radii in group 13 elements is :

   (A) B < Ga < Al < In < Tl

   (B) B < Al < In < Ga < Tl

   (C) B < Ga < Al < Tl < In

   (D) B < Al < Ga < In < Tl

153. In the structure of ClF_3, the number of lone pairs of electrons on central atom 'Cl' is :

   (A) three   (B) one

   (C) four   (D) two

154. The correct order of N-compounds in its decreasing order of oxidation states is :

    (A) NH_4Cl, N_2, NO, HNO_3

    (B) HNO_3, NO, N_2, NH_4Cl

    (C) HNO_3, NH_4Cl, NO, N_2

    (D) HNO_3, NO, NH_4Cl, N_2

155. Which one of the following elements is unable to form MF_6^{3-} ion :

    (A)   In   (B)   Ga

    (C)   B   (D)   Al

156. The compound A on treatment with Na gives B, and with PCl_5 gives C. B and C react together to give diethyl ether. A, B and C are in the order

    (A) C_2H_5OH, C_2H_5ONa, C_2H_5Cl

    (B) C_2H_5OH, C_2H_6, C_2H_5Cl

    (C) C_2H_5Cl, C_2H_6, C_2H_5OH

    (D) C_2H_5OH, C_2H_5Cl, C_2H_5ONa

157. Hydrocarbon (A) reacts with bromine by substitution to form an alkyl bromide which by Wurtz reaction is converted to gaseous hydrocarbon containing less than four carbon atoms. (A) is :

    (A) CH_4 (B) $CH \equiv CH$

    (C) $CH_3 - CH_3$ (D) $CH_2 = CH_2$

158. The compound C_7H_8 undergoes the following reactions:

$$C_7H_8 \xrightarrow{3Cl_2/\Delta} A \xrightarrow{Br_2/Fe} B \xrightarrow{Zn/HCl} C$$
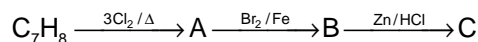
    The product 'C' is :

    (A)   p-bromotoluene

    (B)   m-bromotoluene

    (C)   3-bromo-2, 4, 6-trichlorotoluene

    (D)   o-bromotoluene

159. Which oxide of nitrogen is not a common pollutant introduced into the atmosphere both due to natural and human activity :

    (A)   NO   (B) N_2O_5

    (C) N_2O (D) NO_2

160. The correct difference between first and second order reactions is that

    (A)   the rate of a first-order reaction does depend on reactant concentrations; the rate of a second-order reaction does not depend on reactant concentrations the rate of a first-order reaction does not depend on reactant concentrations

    (B)   the rate of a first-order reaction does not depen on reactant concentrations; the rate of a second - order reaction cannot be catalyzed.

    (C)   a first-order reaction can be catalyzed; a second-order reaction cannot be catalyzed.

    (D)   the half-life of a first-order reaction does not depend on $[A]_0$; the half-life of a second-order reaction does depend on $[A]_0$

161. Among CaH_2, BeH_2, BaH_2 the order of ionic character is :

    (A) $BaH_2 < BeH_2 < CaH_2$ (B) $BeH_2 < CaH_2 < BaH_2$

    (C) $BeH_2 < BaH_2 < CaH_2$ (D) $CaH_2 < BeH_2 < BaH_2$

162. Consider the change in oxidation state of Bromine corresponding to different emf values as shown in the diagram below :

$$BrO_4^- \xrightarrow{1\cdot82\ V} BrO_3^- \xrightarrow{1\cdot5\ V} HBrO$$
$$Br^- \xleftarrow{1\cdot0652\ V} Br_2 \xleftarrow{1\cdot595\ V}$$
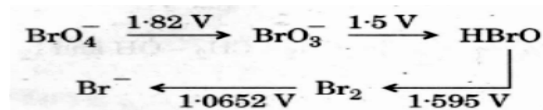
    Then the species undergoing disproportionation is :

    (A) HBrO   (B) BrO_3^-

    (C) Br_2 (D) BrO_4^-

163. In which case is the number of molecules of water maximum :
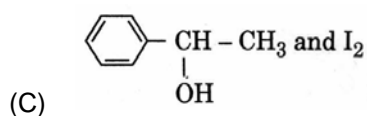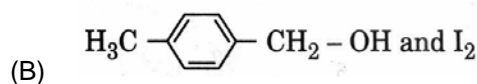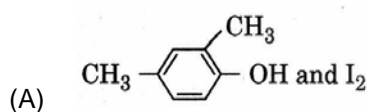
    (A) 10^{-3} mol of water

    (B) 18 mL of water

    (C) 0.00224 L of water vapours at 1 atm and 273 K

    (D) 0.18 g of water

164. Carboxylic acids have higher boiling points than aldehydes, ketones and even alcohols of comparable molecular mass. It is due to their :
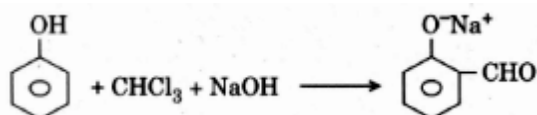
    (A) formation of intermolecular H-bonding

    (B) formation of intramolecular H-bonding

    (C) more extensive association of carboxylic acid via van der Waals force of attraction

    (D) formation of carboxylate ion

165. Compound A, $C_8H_{10}O$, is found to react with NaOI (produced by reacting Y with NaOH and yields a yellow precipitate with characteristic smell.

    A and Y are respectively

    (A)
    CH_3—⟨ring with CH_3, OH⟩ and I_2

    (B)
    H_3C—⟨ring⟩—$CH_2 - OH$ and I_2

    (C)
    ⟨ring⟩—$CH - CH_3$ (with OH) and I_2

    (D)
    ⟨ring⟩—$CH_2 - CH_2 - OH$ and I_2

166. In the reaction

    ⟨ring with OH⟩ $+ CHCl_3 + NaOH \longrightarrow$ ⟨ring with O^-Na^+ and CHO⟩
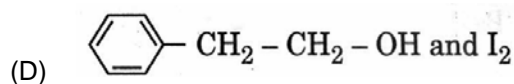
    the electrophilie involved is :

    (A) dichlorocarbene ($:CCl_2$)

    (B) dichloromethyl cation ($\overset{\oplus}{C}HCl_2$)

    (C) dichloromethyl anion ($\overset{\ominus}{C}HCl_2$)

    (D) formyl cation ($\overset{\oplus}{C}HO$)

167. The bond dissociation energies of X_2, Y_2 and XY are in the ratio of 1 : 0.5 : 1. ΔH for the formation of XY is - 200 kJ mol^{-1}. The bond dissociation energy of X_2 will be

    (A) 400 kJ mol^{-1} (B) 200 kJ mol^{-1}

    (C) 800 kJ mol^{-1} (D) 100 kJ mol^{-1}

168. When initial concentration of the reactant is doubled, the half-life period of a zero order reaction :

    (A) remains unchanged   (B) is halved

    (C) is tripled   (D) is doubled

169. The correction factor 'a' to the ideal gas equation corresponds to :

    (A) forces of attraction between the gas molecules

    (B) density of the gas molecules

    (C) electric field present between the gas molecules

    (D) volume of the gas molecules

170. Which one of the following conditions will favour maximum formation of the product in the reaction,

$$A_2\ (g) + B_2\ (g) \rightleftharpoons X_2\ (g),\ \ \Delta_r H = - X\ kJ$$

    (A) High temperature and low pressure

    (B) Low temperature and high pressure

    (C) High temperature and high pressure

    (D) Low temperature and low pressure

171. For the redox reaction

$$MnO_4^- + C_2O_4^{2-} + H^+ \longrightarrow Mn^{2+} + CO_2 + H_2O$$
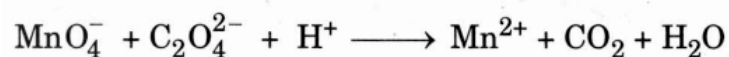
    the correct coefficients of the reactants for the balanced equation are :

	MnO_4^-	$C_2O_4^{2-}$	H^+
(A)	5	16	2
(B)	16	5	2
(C)	2	16	5
(D)	2	5	16

172. Regarding cross-linked or network polymers, which of the following statements is incorrect?

    (A) They contain strong covalent bonds in their polymer chains.

    (B) They contain covalent bonds between various linear polymer chains.

    (C) Examples are bakelite and melamine.

    (D) They are formed from bi- and tri-functional monomers.

173. Nitration of aniline in strong acidic medium also gives m-nitroaniline because :

    (A) In acidic (strong) medium aniline is present as anilinium ion.

    (B) In spite of substituents nitro group always goes to only m-position.

    (C) In absence of substituents nitro group always goes to m-position.

    (D) In electrophilic substitution reactions amino group is meta directive.

174. Which of the following oxides is most acidic in nature :

    (A) CaO   (B) MgO

    (C) BaO   (D) BeO

175. The difference between amylose and amylopectin is :

    (A) Amylose is made up of glucose and galactose

    (B) Amylopectin have $1 \rightarrow 4$ α-linkage and $1 \rightarrow 6$ α-linkage

    (C) Amylopectin have $1 \rightarrow 4$ α-linkage and $1 \rightarrow 6$ β-linkage

    (D) Amylose have $1 \rightarrow 4$ α-linkage and $1 \rightarrow 6$ β-1inkage

176. A 'mixture of 2·3 g formic acid and 4·5 g oxalic acid is treated with cone. H_2SO_4, The evolved gaseous mixture is passed through KOH pellets. Weight (in g) of the remaining product at STP will be -

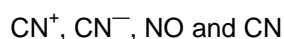   (A) 4·4   (B) 1·4

    (C) 2·8   (D) 3·0

177. Which one is a wrong statement?

    (1) The value of m for $d_z 2$ is zero.

    (2) Total orbital angular momentum of electron in 's' orbital is equal to zero.
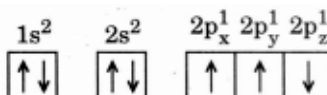
    (C) The electronic configuration of N atom is

    (D) An orbital is designated by three quantum numbers while an electron in an atom is designated by four quantum numbers.

178. Consider the following species :

$$CN^+, CN^-, NO \text{ and } CN$$

    Which one of these will have the highest bond order :

    (A) CN   (B) NO

    (C) CN^+ (D) CN^-

179. Magnesium reacts with an element (X) to form an ionic compound. If the ground state electronic configuration of (X) is .1s2 2s2 2p3, the simplest formula for this compound is

    (A) Mg_3X_2 (B) Mg_2X_3

    (C) Mg_2X (D) MgX_2

180. Iron exhibits bcc structure at room temperature. Above 900°C, it transforms to fee structure. The ratio of density of iron at room temperature to that at 900°C (assuming molar mass and atomic radii of iron remains constant with temperature) is :

    (A) 1/2   (B) $\dfrac{\sqrt{3}}{\sqrt{2}}$

    (C) $\dfrac{3\sqrt{3}}{4\sqrt{2}}$ (D) $\dfrac{4\sqrt{3}}{3\sqrt{2}}$